



# Synthesis of bismuth oxide: Removal of benzene from waters by bismuth oxide nanostructures

Negar Motakef Kazemi<sup>a,\*</sup> and Masoumeh Yaqoubi<sup>b</sup>

<sup>a</sup> Department of Medical Nanotechnology, Faculty of Advanced Sciences and Technology, Tehran Medical Sciences, Islamic Azad University, Tehran, Iran.

<sup>b</sup> Department of Nanochemistry, Faculty of Pharmaceutical Chemistry, Tehran Medical Sciences, Islamic Azad University, Tehran, Iran.

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## ABSTRACT

In this research, the bismuth oxide ( $\text{Bi}_2\text{O}_3$ ) nanostructures were prepared via chemical method at 90 °C for 3 h. The results samples were characterized by Fourier transform infrared (FTIR) for determination of functional groups, X-ray diffraction (XRD) for evaluation of crystal structure, dynamic light scattering (DLS), scanning electron microscope (SEM) for presentation of morphology and size, energy-dispersive X-ray spectroscopy (EDS) for determination of chemical composition, and diffuse reflection spectroscopy (DRS) for ultraviolet (UV) blocking. Also, the  $\text{Bi}_2\text{O}_3$  nanostructures were used for benzene extraction from waters in pH=5-7. By procedure, 30 mg of  $\text{Bi}_2\text{O}_3$  mixed with hydrophobic ionic liquid ([HMIM][PF6]) and injected to water samples. After shaking and centrifuging, benzene removed from water by ionic liquid-micro solid phase extraction (IL- $\mu$ SPE) and determined by gas chromatography with flame ionization detector (GC-FID). The absorption capacity and recovery was obtained 167.8 mg per gram of  $\text{Bi}_2\text{O}_3$  and more than 96%, respectively. Based on the results, the bismuth oxide nanostructures were observed with rod morphology and the diameter of nanometer. The antibacterial activities of the samples were determined against Salmonella using inhibition zone diameter. Based on the study, bismuth oxide nanostructures have good potential for removal of benzene from waters. By IL- $\mu$ SPE method, the results validated by spiking of samples

## 1. Introduction

Nanotechnology has attracted a great attention in very interesting applications in various fields [1, 2]. Today, the nanostructured metal oxides have much attention of researchers actively engaged in various scientific due to their interesting properties and potential applications [3, 4]. One of applied materials is the nanostructures with nanometer-scale rod morphology [5]. Bismuth oxide nanostructure is

one of a good candidate of metal oxide for different applications such as immunosensor [6], gas sensor [7], photocatalyst [8], catalyst [9], preparation of nanostructures [10], photovoltaic [11], biomedical [12, 13], antibacterial effect [14], X-ray shielding [15], optical properties [16], white-light LEDs [17], and magnetic properties [18]. The common methods of bismuth nanostructures synthesis include solution [19], solution combustion [20], solvothermal [21], hydrothermal [22], laser ablation [23], green synthesis [24], sol-gel [25], flame spray pyrolysis [26], thermal decomposition [27], vapor phase deposition [28], and sputtering

\*Corresponding Author: Negar Motakef Kazemi

Email: [motakef@iaups.ac.ir](mailto:motakef@iaups.ac.ir)

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deposition [29]. There are different polymorphs of bismuth oxide including  $\alpha$ - $\text{Bi}_2\text{O}_3$ ,  $\beta$ - $\text{Bi}_2\text{O}_3$ ,  $\gamma$ - $\text{Bi}_2\text{O}_3$ ,  $\delta$ - $\text{Bi}_2\text{O}_3$ ,  $\epsilon$ - $\text{Bi}_2\text{O}_3$ , and  $\omega$ - $\text{Bi}_2\text{O}_3$  which related to the temperature of their formation. The stable polymorph is monoclinic  $\alpha$ - $\text{Bi}_2\text{O}_3$  in low temperature and cubic  $\delta$ - $\text{Bi}_2\text{O}_3$  in high temperature [30]. The increase of temperature was caused the decrease of tetragonal  $\beta$ - $\text{Bi}_2\text{O}_3$  structure and the show monoclinic  $\alpha$ - $\text{Bi}_2\text{O}_3$  in XRD patterns [31]. The ultraviolet light can be caused the increase of risk for skin cancer and ocular damage. The UV radiation included three regions UV-A (320–400 nm), UV-B (280–320 nm), and UV-C (180–280 nm). The earth's atmosphere shields the more harmful UV-C and greater than 99% of UV-B radiation. The UV-A blocking is very important to prevention from hazardous effects of exposure to direct sunlight [32]. Recently, the bismuth oxide was reported as UV-absorber application [33]. Antibacterial activity is another application of bismuth oxide nanoparticles against some pathogenic Gram-negative bacteria [14]. Also, bismuth oxide was used for removal VOCs, BTEX from waters by analytical chemistry. In the present study, bismuth oxide nanostructures were synthesized by chemical method for application of UV blocking and antibacterial activity. In addition, the bismuth oxide ( $\text{Bi}_2\text{O}_3$ ) nanostructures were used for benzene extraction from waters by IL- $\mu$ SPE. Ionic liquid caused to collected solid phase which was extracted by  $\text{Bi}_2\text{O}_3$  in optimized pH.

## 2. Experimental

### 2.1. Materials

All chemicals used were analytical grade. Materials including bismuth nitrate ( $\text{Bi}(\text{NO}_3)_3$ ), nitric acid ( $\text{HNO}_3$ ), and sodium hydroxide ( $\text{NaOH}$ ) were purchased from Merck (Darmshtadt Germany). All aqueous solutions were prepared in deionized water (DW, Millipore). The bismuth oxide ( $\text{Bi}_2\text{O}_3$ ) nanostructures syntheses by Azad University. Benzene (CAS N: 71-43-2;  $\text{C}_6\text{H}_6$ ) and ionic liquid of 1-Hexyl-3-methylimidazolium hexafluorophosphate ([HMIM][PF6]; CAS N:304680-35-1) purchased from Sigma Aldrich.

Five calibration solutions of benzene were prepared and the approximate concentrations of benzene were 0.5, 1.0, 5.0, 10 and 50  $\text{mg L}^{-1}$ . The other chemicals with high purity (99%) were purchased from Sigma (Germany).

For analysis of benzene, gas chromatography based on flame ionization detector (GC-FID) and air sample loop injection (ASL) was used (Netherland). The Agilent 7890A GC can accommodate up to three detectors identified as front detector, back detector, and auxiliary detector. The FID detector chosen was selected for benzene analysis in gas/liquid. Before injection, Slide the plunger carrier down until it is completely over the syringe plunger, and tighten the plunger thumb screw until finger-tight. The injector temperature was adjusted to 200°C and the detector temperature at 250°C. The GC oven temperature was programmed from 30°C to 220°C which was held for 10 min. Hydrogen as the carrier gas was used at a flow rate of 1.0  $\text{mL min}^{-1}$  with split ratio of 1:100.

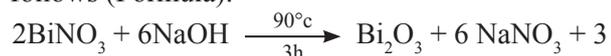
### 2.2. Characterization

The bismuth oxide nanostructures were characterized by Fourier transform infrared spectroscopy, X-ray diffraction, dynamic light scattering, scanning electron microscope, energy-dispersive X-ray spectroscopy, and diffuse reflection spectroscopy. FTIR spectra were recorded on a Shimadzuir 460 spectrometer in a KBr matrix in the range of 400–4000  $\text{cm}^{-1}$ . XRD pattern was performed for evaluation of crystalline structure using a Philips Company X'pert diffractometer utilizing Cu-K $\alpha$  radiation (ASENWARE, AW-XBN300, China). DLS was reported the size and size distribution of nanoparticles (ZEN314, England). SEM was investigated the morphology (KYKY, EM3200, China). EDS was evaluated the elemental and chemical analysis (ASK SEM-CL View VIS, Oxford instruments, UK). DRS was investigated for light absorption and UV blocking properties of nanocomposite (UV2550, Shimadzu).

The antibacterial activity was evaluated using disk diffusion method against *Salmonella* Gram-negative bacteria, strains ATCC 1231, procured from Islamic Azad University.

### 2.3. Synthesis of bismuth oxide nanostructures

Bismuth oxide nanostructures were prepared via chemical method based on schematic reaction as follows (Formula):



In a typical reaction, 0.97 g (0.2 mmol) bismuth nitrate was solved in 1 ml nitric acid, and 9 ml deionized water. Then, 100 ml sodium hydroxide ( $0.1 \text{ mol/L}^{-1}$ ) was added to the resulting solution [8]. The reactants were sealed under reflux and stirred at  $90^\circ\text{C}$  for 3 h. Then, the reaction mixture was cooled to room temperature, and separated by centrifugation. The crystals were washed with deionized water to remove residual salt, and dried in a vacuum oven at  $80^\circ\text{C}$  for 5 h.

### 2.4. Procedure

The bismuth oxide nanostructures based on IL- $\mu$ SPE method was used for extraction of benzene from waters. By procedure, 30 mg of bismuth oxide ( $\text{Bi}_2\text{O}_3$ ) nanostructures mixed with 0.2 g of 1-Hexyl-3-methylimidazolium hexafluorophosphate ([HMIM][PF6]) and diluted with 0.2 mL of pure acetone. Then the mixture was injected into 10 mL of water sample or benzene standard samples with different concentration ( $1\text{-}100 \text{ mg L}^{-1}$ ). The cloudy

solution shaken for 5 min and after centrifuging for 3 min (3500rpm), the upper solution (water or standard solution) was determined by GC-FID. After benzene extraction by IL- $\mu$ SPE procedure, the recoveries of proposed method were measured with the ratio of initial/final concentration of benzene (signal peak area) before determined with GC-FID (Equation EQ1). In addition, adsorption capacity and removal efficiency (RE) was calculated by equation EQ2 and EQ3. A is the initial concentration of benzene in solution and B is final concentration of benzene which determinate by GC-FID in waters. The adsorption capacity of benzene ( $\text{mg g}^{-1}$ ) and, the removal efficiency of benzene (%) was shown in EQ2 and EQ3. The  $C_i$  ( $\text{mg L}^{-1}$ ) and  $C_f$  ( $\text{mg L}^{-1}$ ) are the concentration of benzene before and after extraction procedure,  $V_s$  (L) is the sample volume, and  $M$ (g) is the amount of  $\text{Bi}_2\text{O}_3$ .

$$\text{Recovery} = \frac{A}{B} \times 100 \quad (\text{EQ1})$$

$$AC = \frac{(C_i - C_f) \times V_s}{M} \quad (\text{EQ2})$$

$$RE (\%) = \frac{C_i - C_f}{C_i} \times 100 \quad (\text{EQ3})$$

## 3. Results and Discussion

### 3.1. Fourier transforms infrared

FTIR spectra of bismuth oxide nanostructures were recorded in the range of  $400\text{-}4000 \text{ cm}^{-1}$  with KBr pellets (Fig. 1). The O-H stretching vibrations appear at  $3421 \text{ cm}^{-1}$ . The peak at  $1400 \text{ cm}^{-1}$  is related to C-O vibrations due to organic solvent. The peak at  $435\text{-}505 \text{ cm}^{-1}$  is originated from the metal-oxygen (Bi-O) bond. Fourier transform infrared result is

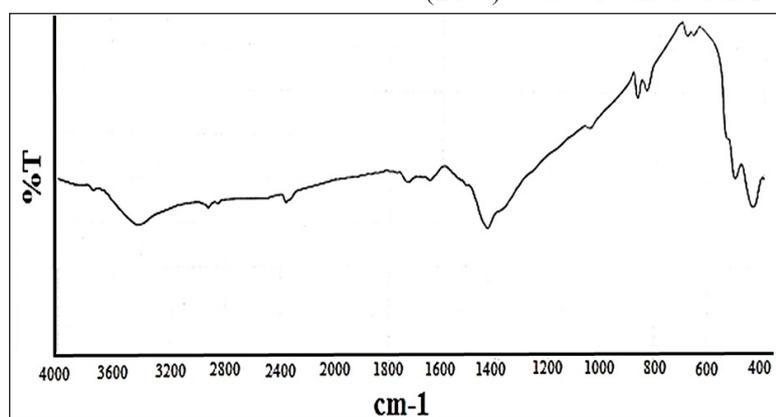


Fig. 1. FTIR bismuth oxide nanostructures.

similar to a previously reported pattern [25].

### 3.2. X-ray diffraction

XRD measurement was used to determine the crystalline structure of bismuth oxide nanostructures in  $2\theta$  range 5 to  $80^\circ$  (Fig. 2). The sharp peak observed at  $2\theta$  around  $28^\circ$ , and all diffraction peaks can be indexed the monoclinic  $\alpha$ - $\text{Bi}_2\text{O}_3$  (JCPDS card No. 41-1449). XRD result is similar to a previously reported pattern [8].

### 3.3. Dynamic light scattering

The dynamic light scattering was used to find out the size and distribution diagram of nanoparticles (Fig. 3). DLS results presented two peaks at 900 nm and  $17 \mu\text{m}$  with narrow distribution at room temperature. The observation of two peaks confirms that the nanostructure is rod shaped.

### 3.4. Scanning electron microscope

The size and morphology structures of samples were studied using SEM that shown rod-shaped with an average diameter of 500 nm, and the length of  $11 \mu\text{m}$  (Fig. 4). SEM result confirmed the DLS result.

### 3.5. Energy-dispersive X-ray spectroscopy

EDS was used to evaluate the chemical composition of bismuth oxide nanostructures. This analysis was clearly showed the identification strong peaks

of bismuth (Bi) and oxygen (O) elements. Based on the result the absorption peaks were exhibited at 2.4, 3.2, 10.8, and 11.8 keV, which illustrated a typical absorption of the metallic bismuth. The energy-dispersive X-ray spectroscopy and mapping of bismuth oxide nanostructures were carried out for elemental analysis (Fig. 5).

### 3.6. Diffuse reflection spectroscopy

DRS absorption spectra of bismuth oxide nanostructures showed UV blocking in three Ultraviolet: UV-A, UV-B, and UV-C (Fig 6). Based on DRS result, the absorption peak was observed 90% ultraviolet in range of 200-400 nm. Based on teh result, bismuth oxide nanostructures are good candidates as UV blocking for research development.

### 3.7. Antibacterial activity

The antibacterial activity was measured against *Salmonella* as Gram-negative bacteria by disk diffusion method for bismuth oxide nanostructures with concentration  $0.01 \text{ g}(\text{mL})^{-1}$ . The zone inhibition was examined approximately 8.6 mm. The cell wall of Gram-negative bacterium is composed a thin layer of peptidoglycan surrounded by a membranous structure called the outer membrane. The presence of carboxylic groups causes to the negative charge of bacterial cells at biological pH. The main mechanisms of antibacterial activity are

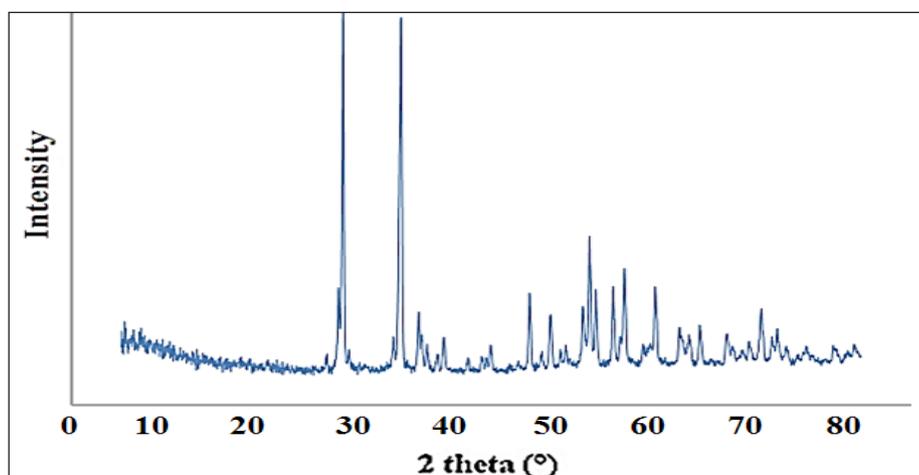


Fig. 2. XRD bismuth oxide nanostructures.

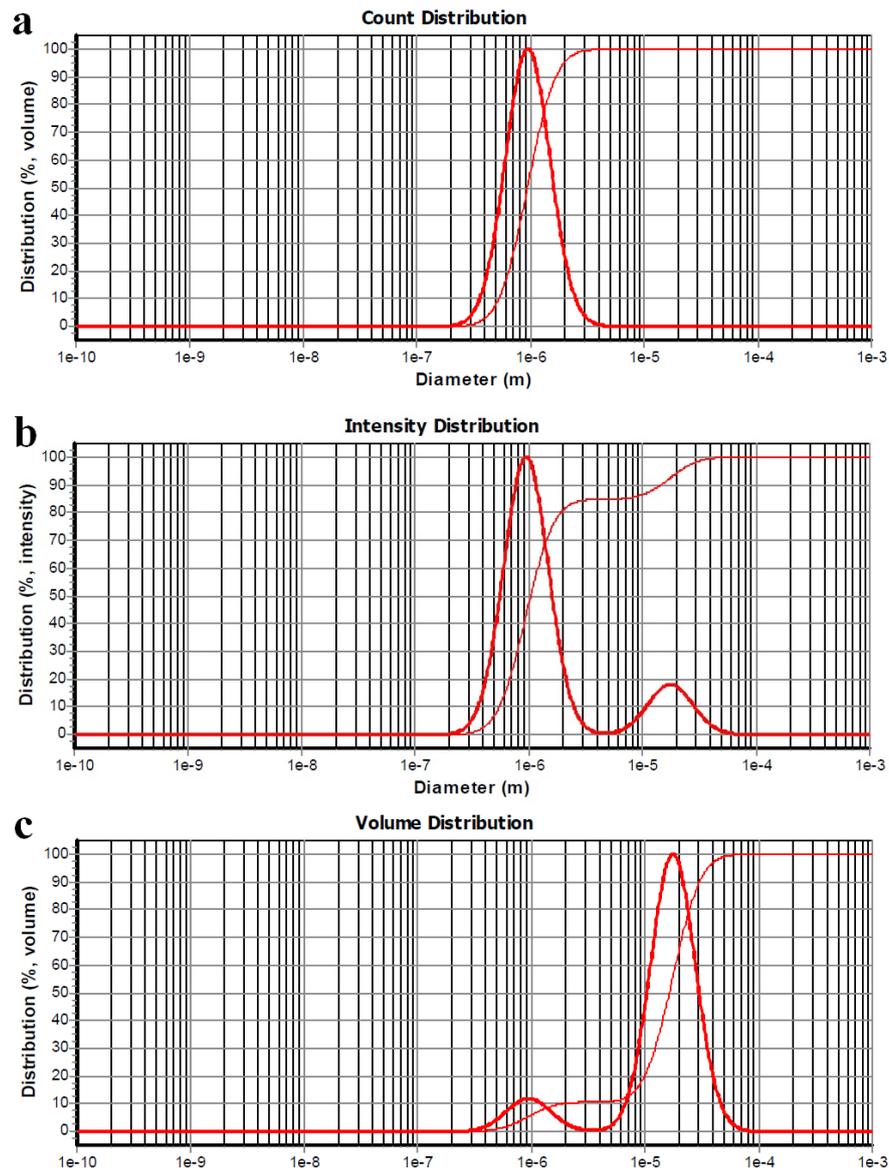


Fig. 3. DLS bismuth oxide nanostructures based on a) count, b) intensity, and c) volume.

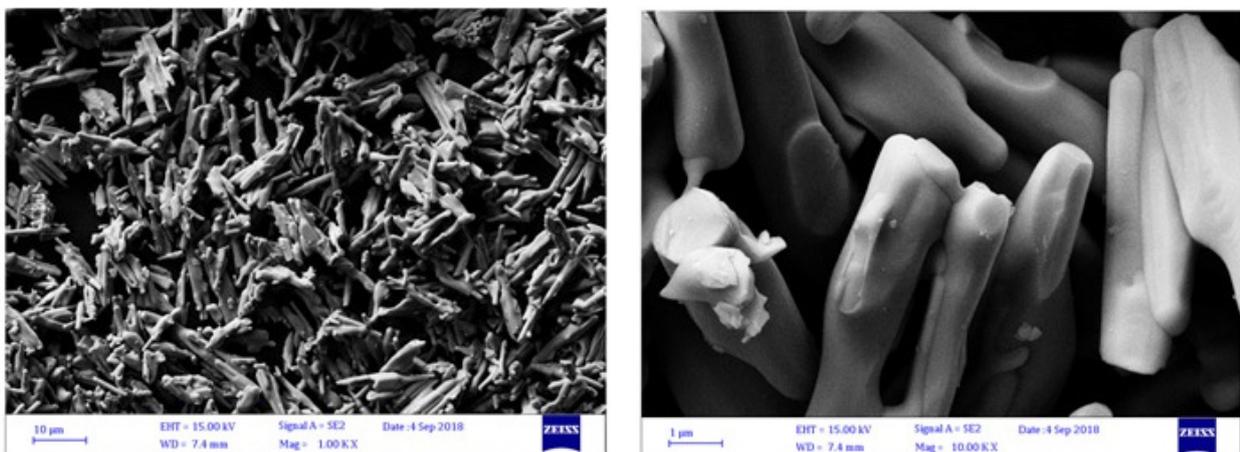
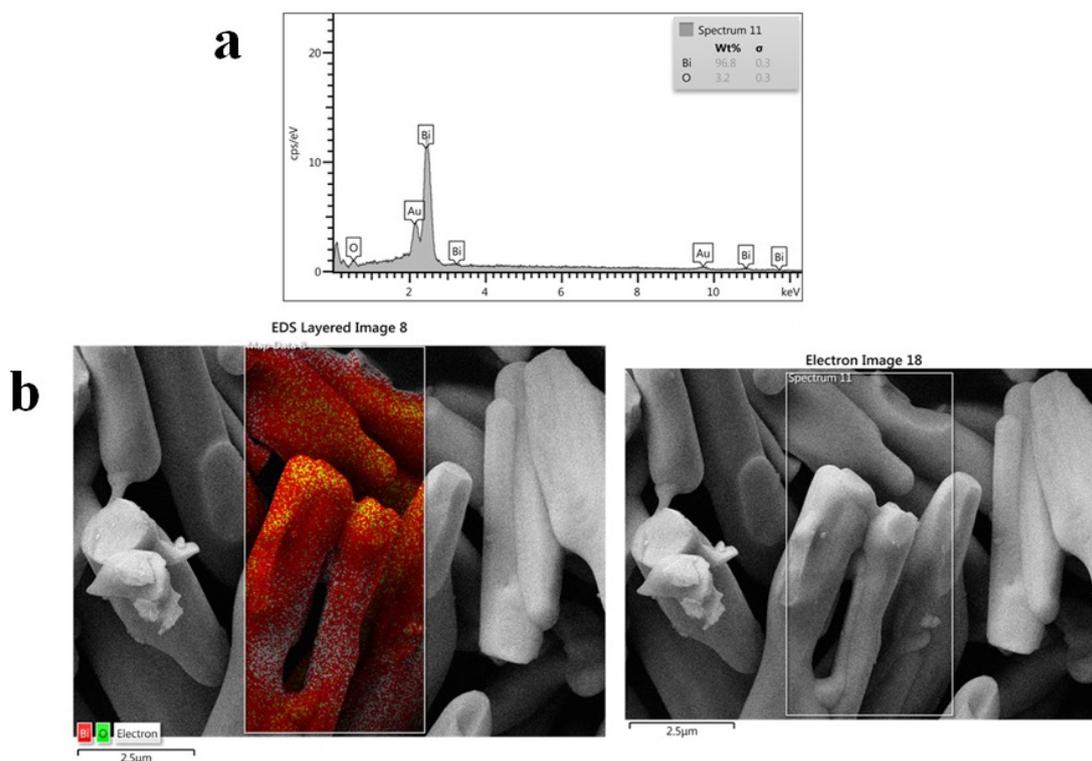


Fig. 4. SEM bismuth oxide nanostructures in different scale bar.

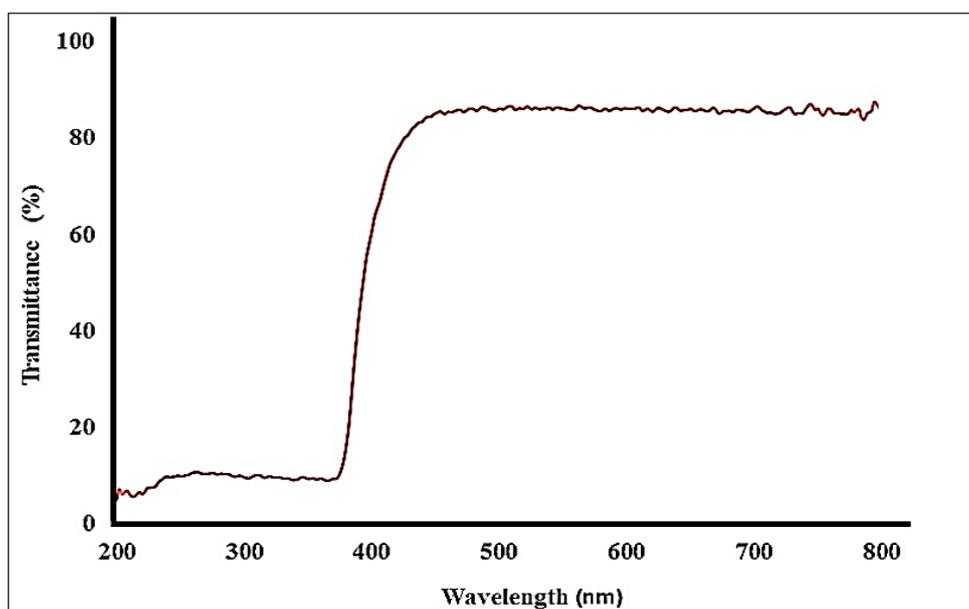
electrostatic forces and adhesion of the opposite charges of Gram-negative bacterium and bismuth oxide nanostructures. Based on the result, bismuth oxide nanostructures are good candidates as antibacterial activity for research development.

### 3.8. Optimizing and validation

The IL- $\mu$ SPE procedure based on the bismuth oxide ( $\text{Bi}_2\text{O}_3$ ) nanostructures was used for extraction of benzene in water and wastewater samples. For increasing of efficient recoveries, all parameters such as, pH, sorbent mass, sample volume, adsorption capacity were studied and optimized. The pH of water sample has a main role for adsorption



**Fig. 5.** a) EDS, and b) elemental map image of bismuth oxide nanostructures.



**Fig. 6.** DRS bismuth oxide nanostructures.

of benzene in water and wastewater by IL-  $\text{Bi}_2\text{O}_3$  by IL- $\mu\text{SPE}$ . The effect of pH range (2-11) on the extraction of benzene was studied containing  $1 \text{ mg L}^{-1}$  and  $10 \text{ mg L}^{-1}$  of  $\text{C}_6\text{H}_6$ . The results showed, the recovery of extraction for benzene were decreased at pH ranges ( $7 < \text{pH} < 5$ ) So, pH of 5-7 was selected as optimized pH for benzene extraction in waters (Fig. 7). By IL- $\mu\text{SPE}$  method, the amount of  $\text{Bi}_2\text{O}_3$  and IL was optimized for extraction benzene in water and wastewater samples. Therefore, 5-50 mg of  $\text{Bi}_2\text{O}_3$  and 0.05-0.4 g of IL was used and optimized. Based on results, more than 25 mg  $\text{Bi}_2\text{O}_3$  and 0.15 g of IL can be extracted benzene in water samples in optimized pH. So, 30 mg and 0.2

g was selected as optimum value for  $\text{Bi}_2\text{O}_3$  and IL, respectively (Fig.8). The sample volume effected on the extraction recoveries of benzene in water samples at pH=5-7. The different sample volumes from 1 to 20 mL ( $1\text{-}10 \text{ mgL}^{-1}$  benzene) were used benzene extraction in water samples by IL- $\mu\text{SPE}$  procedure. The results had good recoveries less than 15 mL of water samples. Therefore, 10 mL was used as the optimal sample volume by proposed procedure (Fig.9).

#### 4. Conclusions

The bismuth oxide nanostructures synthesized by chemical method. The formation of nanostructures

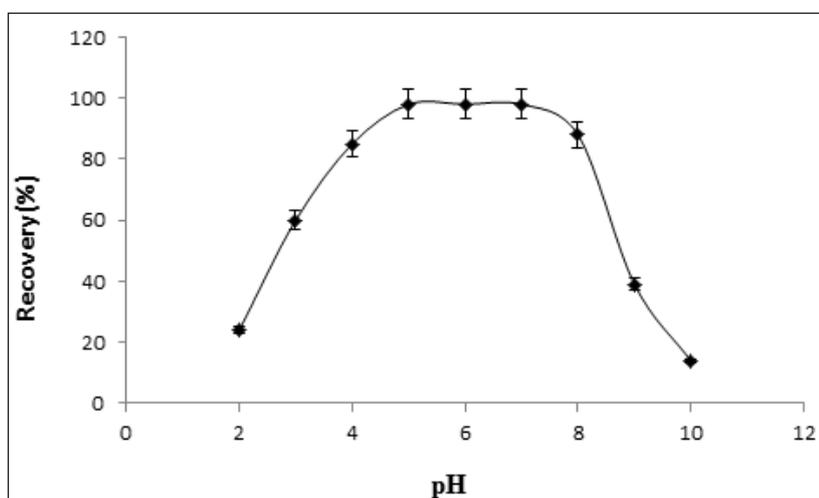


Fig. 7. The effect of pH on benzene extraction in water sample by IL- $\mu\text{SPE}$

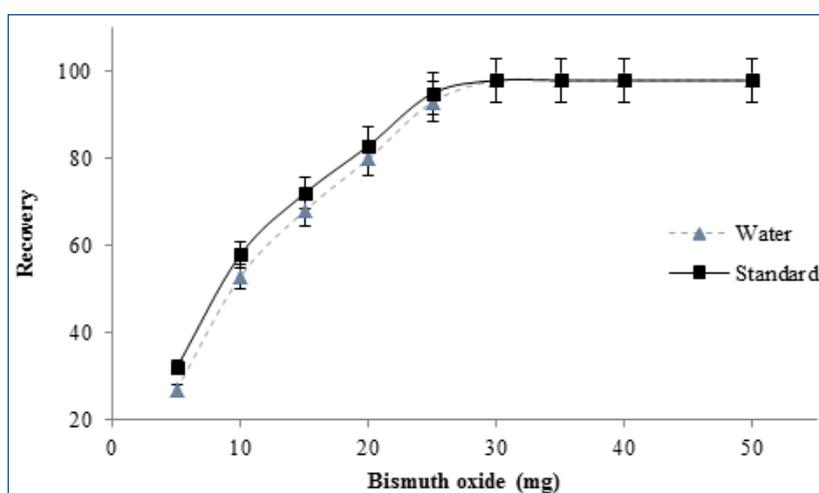


Fig. 8. The effect of bismuth oxide on benzene extraction in water sample by IL- $\mu\text{SPE}$

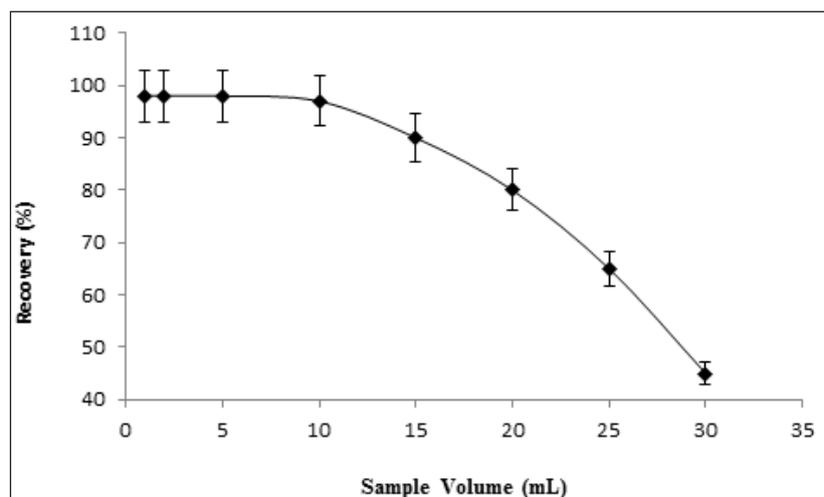


Fig. 9. The effect of Sample volume on benzene extraction in water sample by IL- $\mu$ SPE

was emphasized by DLS with narrow distribution and SEM with rod morphology. XRD confirmed the monoclinic  $\alpha$ - $\text{Bi}_2\text{O}_3$  crystalline structure for bismuth oxide nanostructures. In this study, benzene was extracted from water samples based on IL-  $\text{Bi}_2\text{O}_3$  by IL- $\mu$ SPE procedure at pH=5-7. The absorption capacity and mean of extraction efficiency for  $\text{Bi}_2\text{O}_3$  was obtained 167.8 mg per gram and almost 96%, respectively. Also, we successfully observed UV blocking, and antibacterial activity applications of bismuth oxide nanostructures. These properties can be resulted to many advantages in the future with more safety and less toxicity to human health.

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